

GENERAL CHEMISTRY

STANDARD 10.2

10.2: Describe how changing the pressure, temperature, or surface area of a solution/solute will affect the rate of dissolution

FACTORS THAT AFFECT SOLUBILITY

- Surface area of solute
 - As the surface area of the solute increases, the rate of dissolution increases
 - Reason: There are more points of contact between solute and solvent molecules
- Temperature of the solution
 - As the temperature of the solution increases, the rate of dissolution increases
 - Reason: Higher temperature means molecules are moving faster, increasing the likelihood of a collision between solute and solvent molecules
- Pressure of the solution
 - As the pressure of the solution increases, the rate of dissolution increases
 - Reason: Higher pressure means molecules are packed more tightly together, increasing the likelihood of a collision between solute and solvent molecules

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