GENERAL CHEMISTRY STANDARD 12.2

12.2: Classify a compound as a Bronsted-Lowry Acid or Base

DEFINITIONS

- **Bronsted-Lowry Acid:** Any compound that can donate a proton, or hydrogen ion (H^+), to another compound
- **Bronsted-Lowry Base:** Any compound that can accept a proton from another compound

EXAMPLES

- Examples of Bronsted-Lowry Acids:
  - HCl
  - H_2SO_4
  - HNO_3
  - HBr
- Examples of Bronsted-Lowry Bases:
  - NaOH
  - LiOH
  - $Ca(OH)_2$
  - NH_3

TRY IT YOURSELF

- Identify the following compounds as Bronsted-Lowry Acids or Bronsted-Lowry Bases:
  - $Al(OH)_3$
  - KOH
  - HNO_3
  - HBr
  - $Mg(OH)_2$
  - HNO_3
  - HI
  - NH_3

TRY IT YOURSELF SOLUTIONS

- Identify the following compounds as Arrhenius Acids or Arrhenius Bases:
  - $Al(OH)_3$ **Bronsted-Lowry Base**
  - KOH **Bronsted-Lowry Base**
  - HNO_3 **Bronsted-Lowry Acid**
  - HBr **Bronsted-Lowry Acid**
  - $Mg(OH)_2$ **Bronsted-Lowry Base**
  - HNO_3 **Bronsted-Lowry Acid**
  - HI **Bronsted-Lowry Acid**
  - NH_3 **Bronsted-Lowry Base**